

Experimental and Theoretical Study of the Structure and Reactivity of $\text{Fe}_{1-2}\text{O}_{\leq 6}^-$ Clusters with CO

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Synergistic studies employing experiments in the gas phase and theoretical first principles calculations have been carried out to investigate the structure, stability, and reactivity toward CO of iron oxide cluster anions, Fe_xO_y^- ($x = 1-2$, $y \leq 6$). Collision-induced dissociation studies of iron oxide species, employing xenon collision gas, show that FeO_3^- and FeO_2^- are the stable building blocks of the larger iron oxide clusters. Theoretical calculations show that the fragmentation patterns leading to the production of O or FeO_n fragments are governed both by the energetics of the overall process as well as the number of intermediate states and the changes in spin multiplicity. Mass-selected experiments identified oxygen atom transfer to CO as the dominant reaction pathway for most anionic iron oxide clusters. A theoretical analysis of the molecular level pathways has been carried out to highlight the role of energetics as well as the spin states of the intermediates on the oxidation reaction.

Introduction

Transition metal catalysts, specifically those composed of iron nanoparticles, have been employed in many industrial and biological processes as well as pollution abatement applications.¹ Lin et al. observed nanosized iron oxides to be highly active for CO oxidation at low temperatures.² In addition to other studies probing the oxidation of carbon monoxide with iron,³ iron oxides showed activity for the oxidation of methane⁴ and various hydrocarbons.⁵ Iron oxides are also promising because they exhibit good catalytic lifetimes and resistance to high concentrations of moisture and CO_2 , which often poison catalysts.⁶ Therefore, iron oxides may become economical alternatives to costly precious metal catalysts.

Studies involving iron and iron oxide have been conducted previously through both calculations⁷⁻⁹ and photoelectron spectroscopy experiments.¹⁰⁻¹² FeO , Fe_2O_3 , and Fe_3O_4 are stable stoichiometries in the condensed phase; thus, they are the most studied species.^{10,13} It was suggested by Li and co-workers that Fe_2O_3 was the most active species for oxidation, showing that Fe_2O_3 behaved as a catalyst in the presence of O_2 and as a reagent in the absence of oxygen, oxidizing CO to CO_2 .¹³ Sequential reduction of Fe_2O_3 in the absence of O_2 produced Fe_3O_4 , FeO , and Fe species. Previous calculations of the reaction of neutral Fe_2O_3 with CO led to a proposed mechanism with a viable energetic pathway to produce CO_2 .¹⁴⁻¹⁶ Through cooperative effects, the adsorption of CO onto the cluster was predicted to weaken a Fe–O bond and made this oxygen readily available to a second CO for easy formation of CO_2 .

Iron oxide layers are commonly employed in catalytic supports and have shown active participation in the reactions.^{6,17} Iron is unique and more active than other supports because it is

easily reduced and allows anion vacancy formation due to a highly disordered structure.^{18,19} Anion vacancies are important in the catalytic process because they stabilize the supported clusters and influence charge density. Schubert et al. studied several different transition metal supports and found that iron oxides were able to adsorb large amounts of oxygen.¹⁷ In addition to the adsorbed oxygen participating in the process, they proposed a mechanism that allowed for the oxide support to supply the oxygen for the reaction. The ability of iron oxide to function both ways greatly enhanced the oxidation activity.

Gas-phase cluster studies are especially amenable to elucidating the structure and active sites of reactive species. These studies eliminate the uncertainty of different catalyst preparation methods and allow for fundamental insight into site specific activity. Clusters, therefore, are ideal models to identify what occurs on a catalytic surface and to help elucidate molecular level reaction mechanisms. Numerous studies have shown similar products from reactions of metal oxides in both the condensed and gas phase.²⁰⁻²⁴ Our present joint experimental and theoretical study aids in uncovering species with increased activity and selectivity for the oxidation of CO to CO_2 . The results provide important insight into the nature of the active sites responsible for condensed phase catalysis, shedding light on the role of cluster size, charge and oxidation state, composition, and stoichiometry toward CO oxidation in the presence of iron oxide. We focus on anionic iron oxide behavior herein, and address charge state and electron density effects on the structure and bonding of small cationic iron oxide clusters in a separate publication.²⁵

Experimental Methods

Gas-phase cluster studies were performed using a guided ion beam mass spectrometer coupled to a laser vaporization source, explained in detail previously.²⁶ Briefly, the second harmonic of a Nd:YAG laser was used to ablate a rotating and translating iron rod (PVD Materials Corp., 99.95% purity). At a predeter-

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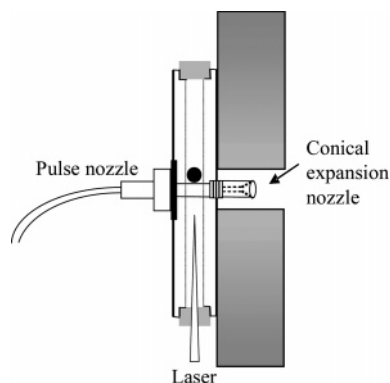


Figure 1. Laser vaporization source showing conical expansion nozzle screwed into the exit channel.

mined time, oxygen seeded in helium (~1%) was pulsed over the rod, forming a dense, hot plasma. A 27 millimeter (mm) conical expansion nozzle was placed at the exit of the source, shown schematically in Figure 1, allowing for more third-body collisions and aiding in larger cluster formation. The nozzle was composed of a 15 mm long channel with a 4 mm diameter and 12 mm conical expansion cone with a 30° total internal angle. The clusters then underwent supersonic expansion in a field free region before passing through a 3 mm skimmer which created a molecular beam. The cooled clusters were then focused by a set of electrostatic lenses and deflectors into the first quadrupole mass filter. Each cluster species was individually selected in the first quadrupole and directed through a second set of lenses into the octopole reaction cell. Carbon monoxide reactant gas, ranging from 0 to 20 mTorr, was added to the reaction cell and the pressure was monitored by a MKS baratron. The products were focused by a third set of lenses into the second quadrupole where they were mass analyzed and finally detected using a channel electron multiplier. Studies were also conducted with nitrogen in the reaction cell under the same conditions of pressure and energy for verification of a chemical reaction with carbon monoxide. Since both CO and N₂ are of the same nominal mass, experiments with N₂ aided in identifying collisional fragmentation products. Collision-induced dissociation (CID) experiments were conducted to study the fragmentation patterns of the iron oxide clusters. In these experiments, inert xenon gas was introduced into the reaction cell under single collision conditions (0.09 mTorr) while the kinetic energy of the ions in the octopole reaction cell was slowly raised from 0 to 40 eV laboratory frame energy. Experiments with slightly higher collision pressures, 0.2 mTorr of Xe, were conducted to find sequential fragmentation patterns under multiple collision conditions.

Theoretical Methods

The theoretical studies were carried out using two different numerical schemes that were developed within a density functional formalism.²⁷ The exchange and correlation effects were incorporated through the generalized gradient approximation (GGA) via a functional proposed by Perdew, Burke, and Ernzerhof.²⁸ The electronic structure was determined using a linear combination of atomic orbitals molecular orbital approach. The wave function for the cluster was constructed by a linear combination of Gaussian type orbitals centered at the atomic positions in the cluster. The actual calculations employed two different numerical programs. Most calculations were carried out using the Naval Research Laboratory Molecular Orbital Library (NRLMOL) set of codes developed by Pederson and

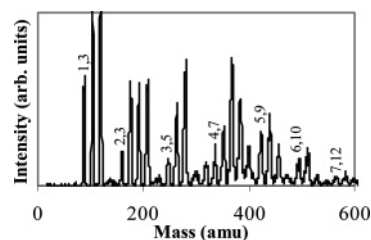


Figure 2. A typical mass distribution produced for iron oxide anionic clusters. The first iron oxide in each series is labeled according to Fe_xO_y⁻, where (x,y) with subsequent peaks in the series have one additional oxygen atom.

co-workers.²⁹⁻³¹ For these calculations, we employed a 5s, 4p, and 3d basis set for the C and O atoms and 7s, 5p, and 4d basis for the Fe atom.³¹ In each case, the basis set was supplemented by a diffuse Gaussian. For more details, the reader is referred to the original papers.²⁹⁻³¹ Supplementary calculations were also carried out using the deMon2K software³² in order to eliminate any uncertainties associated with the choice of basis set or the numerical procedure. In these studies, we employed a gradient-corrected density functional²⁸ and the double- ζ valence polarized (DZVP) basis sets³³ for C and O and the Wachters-F basis set³⁴ for Fe. The GEN-A2 auxiliary function set for C and O and the GEN-A2* auxiliary function set for Fe were used. For each cluster structure, the configuration space was sampled by starting from several initial configurations. Then the geometry was optimized by moving the atoms in the direction of the forces until they dropped below a threshold value. Since transition metal atoms are marked by nonzero spin multiplicities, the calculations included optimizing the spin multiplicities of each cluster.

Results and Discussion

The anionic iron oxide cluster distribution with both dissociated and molecular oxygen adsorbed at near thermal energy is shown in Figure 2. Expansion gas mixtures ranging from 1–20% oxygen were investigated, with the lowest percentage of oxygen in helium providing the broadest cluster distribution. Various conditions at the exit of the source were also examined and experiments conducted with no conical expansion nozzle produced only monomer iron oxide species. A conical nozzle placed at the end of the source region before the field free expansion region (see Figure 1) allowed for more collisions and thus better iron oxide clustering. A 27 mm nozzle was used because the longer nozzle (51 mm) produced species that were more oxygen rich possessing large numbers of molecularly bound O₂ units. We did not generate any stoichiometric Fe_xO_y⁻ clusters with $x = y$ having the same number of iron and oxygen atoms. It is interesting to note that Bernstein and co-workers produced neutral iron oxide clusters with oxygen deficient stoichiometries by varying the oxygen concentration in the system.³⁵ They made oxygen deficient iron oxide clusters employing a 0.1% oxygen mixture. However, in our source this concentration produced a narrower mass distribution of iron oxide clusters than the 1% oxygen mixture. The following studies include FeO₂₋₄⁻ and Fe₂O₃₋₆⁻ anionic species, mostly with an oxygen-rich metal to oxygen ratio of $x/y < 1$.

Structures. Figure 3 shows the calculated optimized lowest-energy structures for the FeO₁₋₄⁻ and Fe₂O₂₋₆⁻ clusters. For iron oxide clusters containing a single Fe atom, oxygen atoms bind directly to the metal with no molecular oxygen units. The maximum coordination number of Fe was four in the tetrahedral form of FeO₄⁻. For FeO₅⁻ (not shown), the fifth oxygen attaches

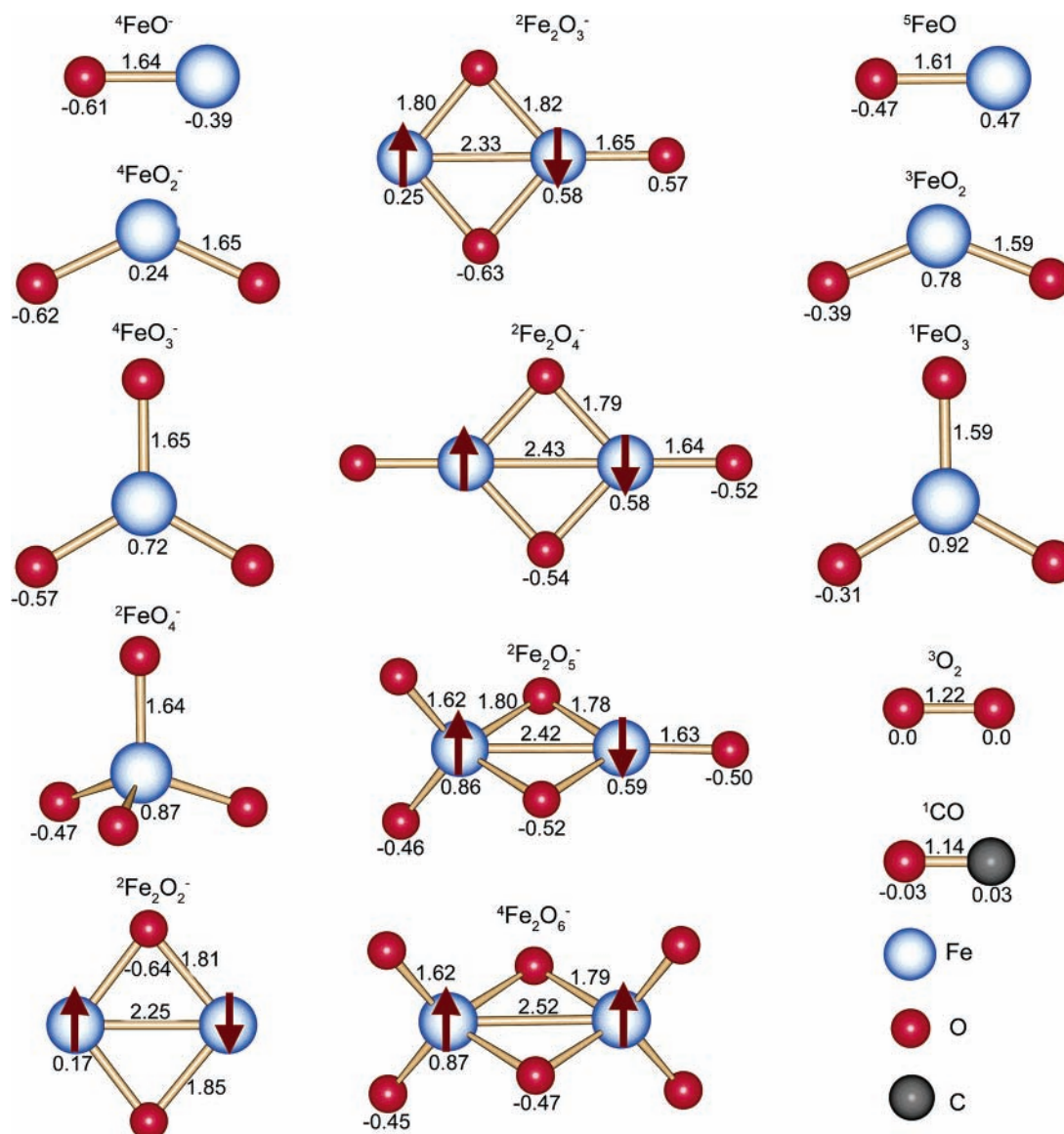


Figure 3. The ground state geometries of O_2 , CO , FeO_{1-3} , and Fe_xO_y^- clusters. The bond lengths are given in Angstroms and the superscripts indicate the spin multiplicity. The arrows indicate the spin polarization at the Fe atoms for the Fe_2O_y^- clusters. The Mulliken charges are marked below each atom.

to an oxygen atom forming an O_2 unit. In the case of FeO_6^- (not shown), a FeO_4^- unit with a O_2 molecule is bound to an oxygen atom at a distance of 3.1 Å. A higher-energy isomer for FeO_4^- possessing a molecular oxygen subunit was 0.78 eV above ground-state energy. The Fe–O bond lengths for FeO_{1-4}^- are relatively similar. The spin multiplicity changes from quartet in FeO_{1-3}^- species to doublet for FeO_4^- .

In the iron oxide clusters containing two Fe atoms, a basic ring structure composed of Fe_2O_2^- with oxygen bridging each iron atom is formed. Fe_2O_3^- and Fe_2O_4^- have their third and fourth oxygen atoms attached to iron outside the ring. The addition of oxygen atoms to the cluster shortens the Fe–O bond length from 1.85 to 1.79 Å within the ring and 1.65 to 1.62 Å outside. Maximum coordination was obtained for the Fe_2O_6^- cluster with two extra oxygen atoms bound directly to each Fe atom. A higher-energy isomer of Fe_2O_6^- with an O_2 subunit was found 1.23 eV above the ground state. An anti-ferromagnetic spin coupling was found for Fe_2O_2^- , Fe_2O_3^- , Fe_2O_4^- , and Fe_2O_5^- cluster species. However, the Fe sites in Fe_2O_6^- were found to be coupled ferromagnetically. Similar progressions of the exchange coupling with oxidation have been reported for the case of Cr_2O_n clusters.³⁶

CID Fragmentation. Collision induced dissociation studies were undertaken with inert xenon gas under single (0.09 mTorr) and multiple (0.2 mTorr) collision conditions and the results are listed in Table 1. The fragmentation of selected clusters and their corresponding neutral loss aids in elucidating the structures of charged clusters and obtaining the general order of bond strength. Trends have been established between the calculated dissociation energies (DE) and the order of experimental fragmentation products, which are discussed below.

For FeO_n^- clusters containing a single Fe atom, the observed fragmentation channels with Xe correspond to the loss of an O atom or O_2 molecule. For FeO_2^- and FeO_3^- , the pathway at the lowest energy corresponds to the loss of an atomic O, while for FeO_4^- , the loss of O_2 is the lowest-energy pathway. At higher collision energies, FeO_4^- does lose an O atom. More interesting are the fragmentation channels for Fe_2O_n^- clusters containing two Fe atoms. In the case of Fe_2O_3^- , the product channel that is observed at the lowest collision energy is loss of an O atom. In contrast, the product channel that is observed at the lowest collision energy for Fe_2O_4^- is loss of FeO . In the case of Fe_2O_5^- , the lowest-energy channel involves the loss of a FeO_2 , while for Fe_2O_6^- , the loss of an O_2 molecule is the

TABLE 1: The Collision Induced Dissociation Fragmentation Channels, Calculated Dissociation Energies (DE), and Reaction Products with CO and N₂ for Fe_{1,2}O₂₋₆⁻ Clusters

Fe _x O _y ⁻ (x, y)	products with Xe ^a	neutral(s) lost ^b	DE (eV)	products with CO	products with N ₂
1, 2	1, 1	0, 1	6.03	1, 1	no products
1, 3	1, 2	0, 1	5.8	1, 2	no products
1, 4	1, 2	0, 2	3.47	1, 2	1, 2 ^c
2, 3	1, 3	0, 1	3.87	1, 3 ^c	
	2, 2	0, 1	6.10	2, 2	no products
	1, 2	1, 1	3.70		
2, 4	1, 3	1, 0	3.41		
	1, 3	1, 1	3.94	2, 3	no products
	1, 2	1, 2	4.68	1, 3	
2, 5	2, 3	0, 1	6.03		
	1, 3	1, 2	3.44	2, 4	no products
	2, 4	0, 1	4.56	1, 3	
2, 6	2, 3	0, 2 ^c	4.39		
	2, 4	0, 2 ^d	2.90	2, 4	2, 4 ≈ 2, 5 ^c
	2, 5	0, 1	4.55	2, 5 ^c	
	1, 3	1, 3 ^c	3.22		

^a Fragmentation channels are shown in order of observation with increasing collision energy. ^b Neutral loss is assigned based on the difference between the selected cluster and fragment ion formed. ^c Channels taking place under multiple collision conditions. ^d Dissociation occurs at near thermal energies.

lowest-energy channel. In each case, other channels do open up as the energy is increased. The basic question is then whether the onset of these pathways may be explained by the energetics of fragmentation.

To answer this question, we calculated the energies required to remove an O, O₂, O⁻, and O₂⁻ from FeO_n⁻ and Fe₂O_n⁻ clusters and the energies for fragmentation pathways leading to the production of Fe, FeO, FeO₂, and FeO₃ in the case of Fe₂O_n⁻ clusters. The results of these investigations are plotted in Figure 4 and the exact values are provided in the supplemental Tables S1, S2, S3, S4, and S5. The calculated dissociation energies for various reaction channels corresponding to observed experimental products are recorded in Table 1. We first discuss the trends in the theoretical results in Figure 4.

Since the current experiments detect anionic species, only the loss of neutral O and O₂ subtracted from the mass selected cluster can be experimentally observed. For FeO_n⁻ clusters, one notices that the energy to remove an O atom or an O₂ molecule decreases with increasing oxygen coverage. As an example, while it takes 5.99 eV to remove an O atom from FeO₂⁻, only 3.86 eV is needed to remove an O atom from FeO₄⁻. This decrease makes the removal of O₂ a favorable channel at higher oxygen coverage particularly since energy may be gained when two O atoms combine to form an O₂ molecule. Indeed, for FeO₄⁻, the energy to remove an O₂ molecule is noticeably less than to remove an O atom, while for FeO₅⁻, it will take only 0.95 eV to remove an O₂ molecule. This is generally consistent with the observed fragmentation patterns in Table 1, which shows the loss of O as the preferred pathway for FeO₂⁻ and FeO₃⁻, while the lowest pathway for FeO₄⁻ fragmentation is the loss of an O₂ molecule.

For Fe₂O_n⁻ clusters, the fragmentation pathways include the production of Fe, FeO, FeO₂ and possibly FeO₃ fragments. The dissociation energies are shown graphically in Figure 4; the exact values are provided in a supplemental Table S1. The dissociation energy values associated with experimental products are also recorded in Table 1. It is interesting to note that the loss of O₂ reaches an energetic minimum for Fe₂O₆⁻ as the energy needed to break two Fe–O bonds is likely to be overcome by the exothermic formation of O₂ releasing 6.21 eV of energy. Oxygen

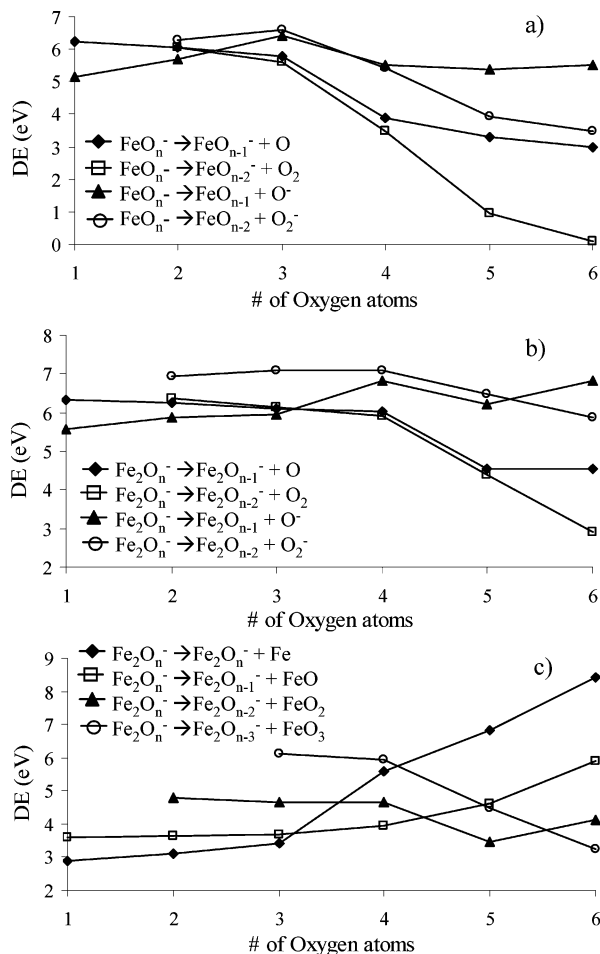


Figure 4. Graphs of the dissociation energy associated with removing an O atom or O₂ subunit from (a) FeO_n⁻ and (b) Fe₂O_n⁻. (c) Graph of the dissociation energy associated with removing Fe or FeO_n from Fe₂O_n⁻ clusters.

recombination was also observed in CID studies previously conducted in our laboratory with V₂O₅⁺.³⁷ The structure of V₂O₅⁺ contained only atomic oxygen bonds to the metal;³⁸ however, dissociation of O₂ was observed.

Breaking of Fe–Fe bonds occurred in the dimer clusters, mostly at higher energies or under multiple collision conditions. The core structure of the dimer clusters after fragmentation was either FeO₃⁻ or FeO₂⁻, which represent the stable building blocks of the larger iron oxide clusters. Fragmentation of neutral FeO from the Fe₂O₄⁻ cluster requires 3.94 eV of overall energy as shown in Figure 5 that depicts various steps in the dissociation process. The dissociation of Fe₂O₄⁻ begins with the breaking of the bond between an Fe atom and a bridging O atom, which changes its coordination to become a terminal O atom, requiring 0.81 eV. Breaking the Fe–Fe bond and internal rotation of an FeO₂ subunit then leads to a chain structure and requires an additional 0.29 eV. From this open structure, it takes an additional 2.84 eV to lose neutral FeO and 3.59 eV to lose FeO₂. The fragmentation of neutral FeO₂ from Fe₂O₅⁻, shown in Figure 6, only requires 2.83 eV after structural rearrangement. Figure 4 shows that more energy is needed to fragment neutral FeO₂ from Fe₂O₄⁻ than neutral FeO; however, this order is reversed in the case of Fe₂O₅⁻. Therefore our theoretical results verify the experiments that show dissociation of FeO from Fe₂O₄⁻ is observed at lower energy while fragmentation of FeO₂ required additional energy. These two clusters are unique because they form the stable fragment FeO₃⁻ prior to oxygen loss. This indicates that the Fe–O bonds for Fe₂O₄⁻ and Fe₂O₅⁻

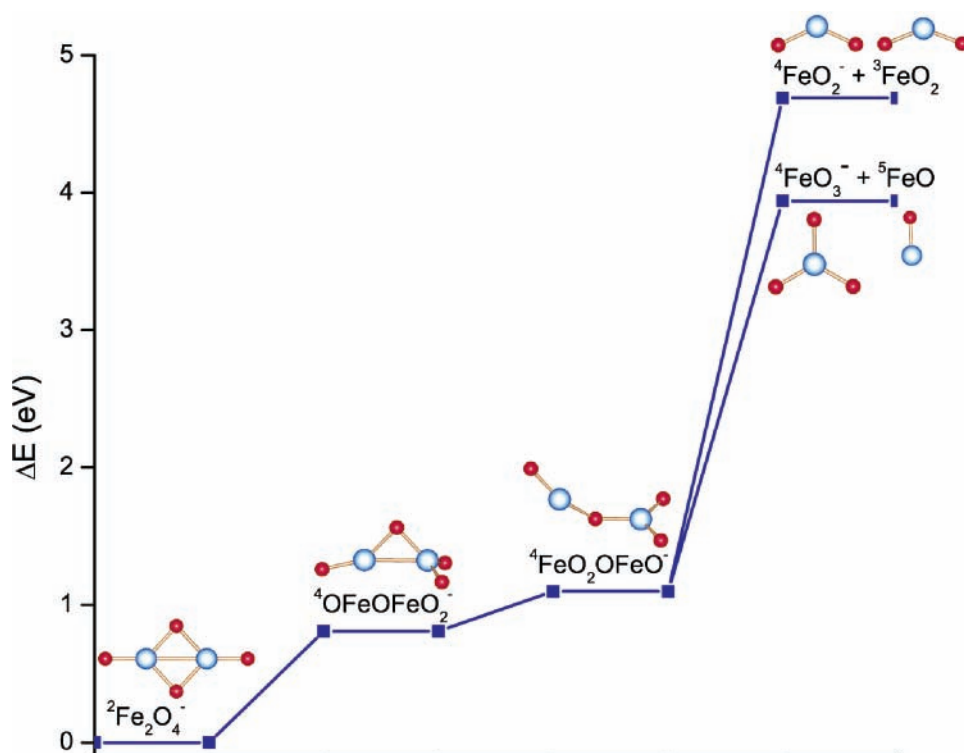


Figure 5. The change in energy (ΔE) for the dissociation pathway of Fe_2O_4^- to FeO_3^- and FeO_2^- . The superscripts indicate the spin multiplicity.

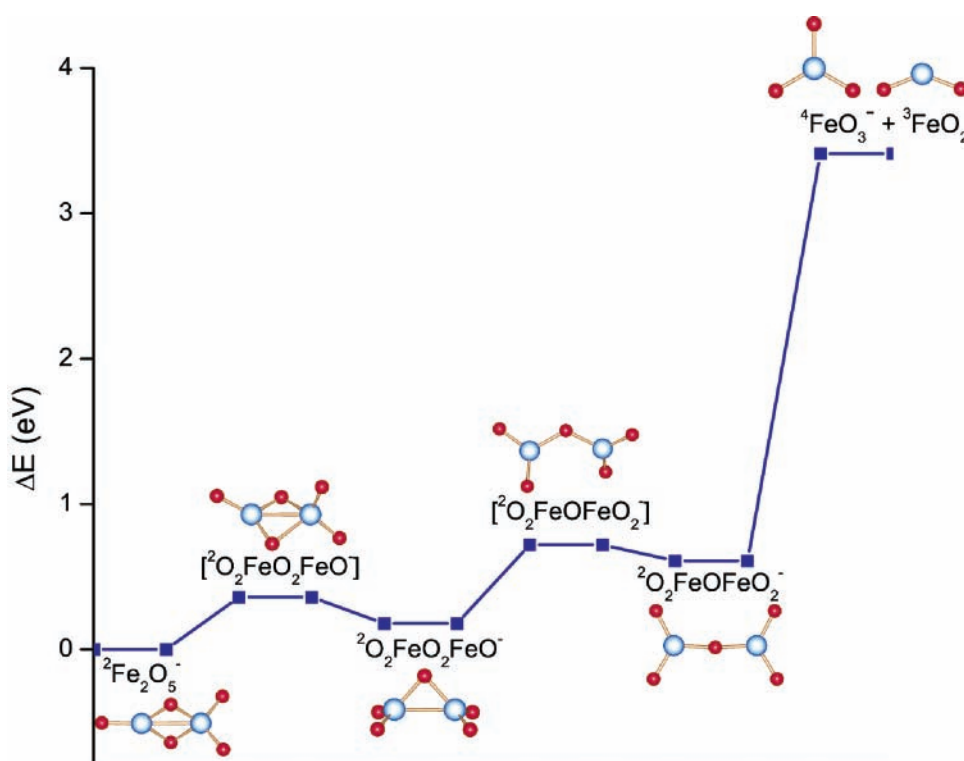


Figure 6. The change in energy (ΔE) for the dissociation pathway of Fe_2O_5^- to FeO_3^- . The superscripts indicate the spin multiplicity.

are very strong. Indeed, it requires 6.03 and 4.56 eV to break the Fe–O bonds, respectively.

In the case of Fe_2O_3^- , loss of an oxygen atom is the first observed fragment. This fragment does not follow the pattern of dissociation energies listed in Table 1. Figure 7 shows the change in energy for the dissociation pathways for neutral Fe and FeO loss. Supplemental Table S4 shows the corresponding energies for each reaction step associated with the two possible bond fragmentations. In both pathways, multiple bond frag-

mentations are required for neutral Fe and FeO loss. In the first step, a bridging Fe–O bond is broken and the oxygen either bonds to the Fe atom that is less coordinated requiring 0.59 eV or stays on the iron atom with another oxygen atom already attached requiring 0.90 eV of energy. At this stage in the reaction path, spin is conserved in the doublet state for the neutral FeO dissociation pathway. From the $[\text{}^2\text{O}_2\text{FeOFeO}_2]$ transition state, the bonds are relaxed and followed by Fe–Fe dissociation to form a linear chain of alternating Fe–O bonds. Finally, neutral FeO

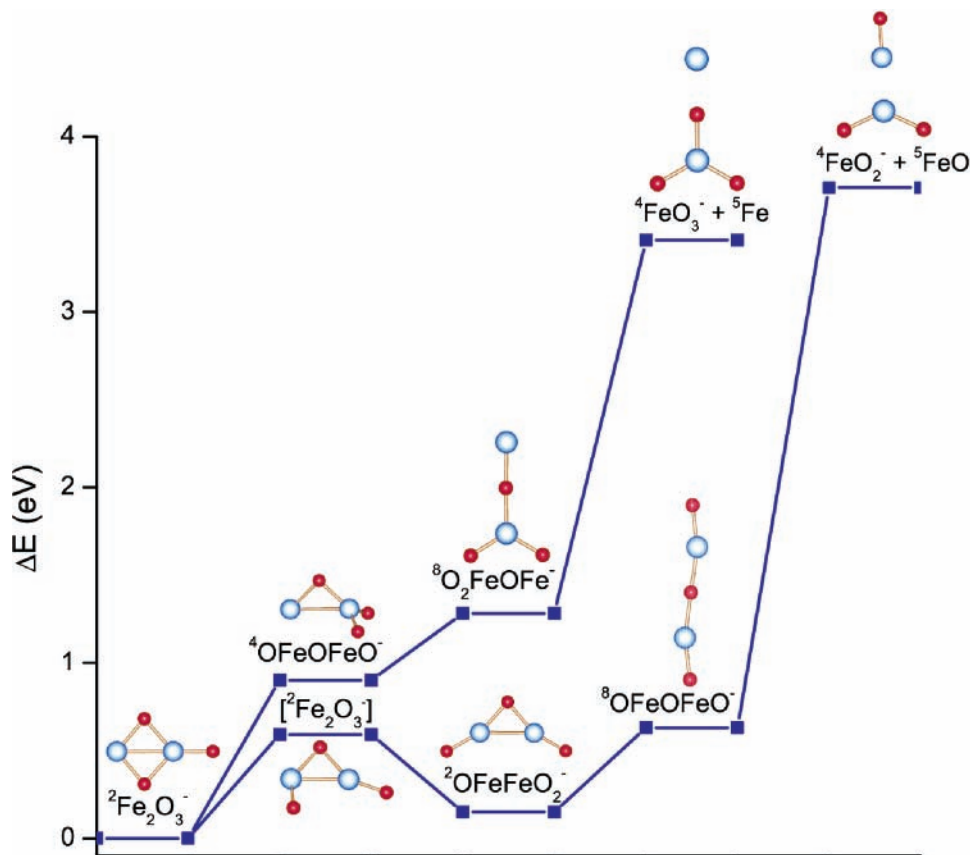


Figure 7. The change in energy (ΔE) for the dissociation pathway of Fe_2O_3^- to FeO_3^- and FeO_2^- . The superscripts indicate the spin multiplicity.

is released with an additional 3.08 eV of energy required. The other pathway for neutral Fe dissociation also requires a spin change and bond rearrangements. Subsequent neutral Fe loss from the cluster requires an additional 2.14 eV of energy.

While a comparison of the two pathways may indicate that the loss of Fe is energetically more favorable, the pathway for the loss of FeO involves more stable intermediate steps. We believe that this accounts for the FeO fragments being observed before Fe loss. Further, the numerous bond breaking steps and rearrangement of the atoms along the pathways may make these FeO and Fe loss mechanisms kinetically unfavorable compared to the energetically unfavorable O atom loss channel. In both cases (loss of FeO and loss of Fe), the need to undergo a spin transition from a doublet state to an octet intermediate state is expected to slow the reaction considerably.³⁹

Additionally the calculated dissociation pathway for Fe_2O_6^- , which support our CID results, are provided in the Supporting Information (Figure S1, Table S5). For Fe_2O_6^- the fragmentation order observed with CID differs from the calculated dissociation energies. For the same reasons discussed above for Fe_2O_3^- , FeO_3 loss occurs after removal of an O atom. Multiple bond fragmentation in Fe_2O_6^- requires multiple collisions and therefore is observed after O atom loss.

Reactions with CO and N₂. The reaction products for each cluster species are recorded in descending order of relative product intensity in Table 1. Theoretical studies indicate that the atomization energy of CO and CO₂ are 11.63 and 17.97 eV, respectively. The formation of CO₂ is, therefore, energetically feasible in cases where it takes less than 6.34 eV to remove an O atom from the cluster. However, the presence of reaction barriers can prevent the formation of CO₂. This is the reason that CO₂ formation does not occur in the presence of O₂ even

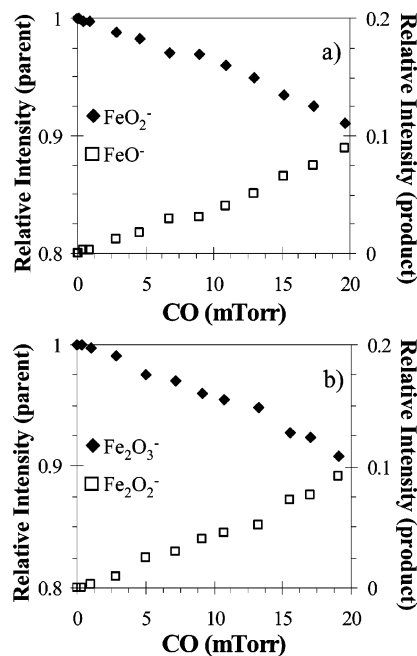


Figure 8. Relative intensities of (a) FeO_2^- and FeO^- , and (b) Fe_2O_3^- and Fe_2O_2^- as functions of CO reactant gas pressure. The observed behavior shows that both FeO_2^- and Fe_2O_3^- are effective for the oxygen atom transfer reaction.

though the binding energy of O₂ is only 6.20 eV. As Table 1 shows, the energy to remove an O from the cluster species, FeO_2^- , FeO_3^- , FeO_4^- , Fe_2O_3^- , Fe_2O_4^- , and Fe_2O_5^- is less than 6.34 eV, and they are all active toward oxygen atom transfer to CO. This process occurs as a dominant reaction channel for most anionic species. Since neutral species are not detected,

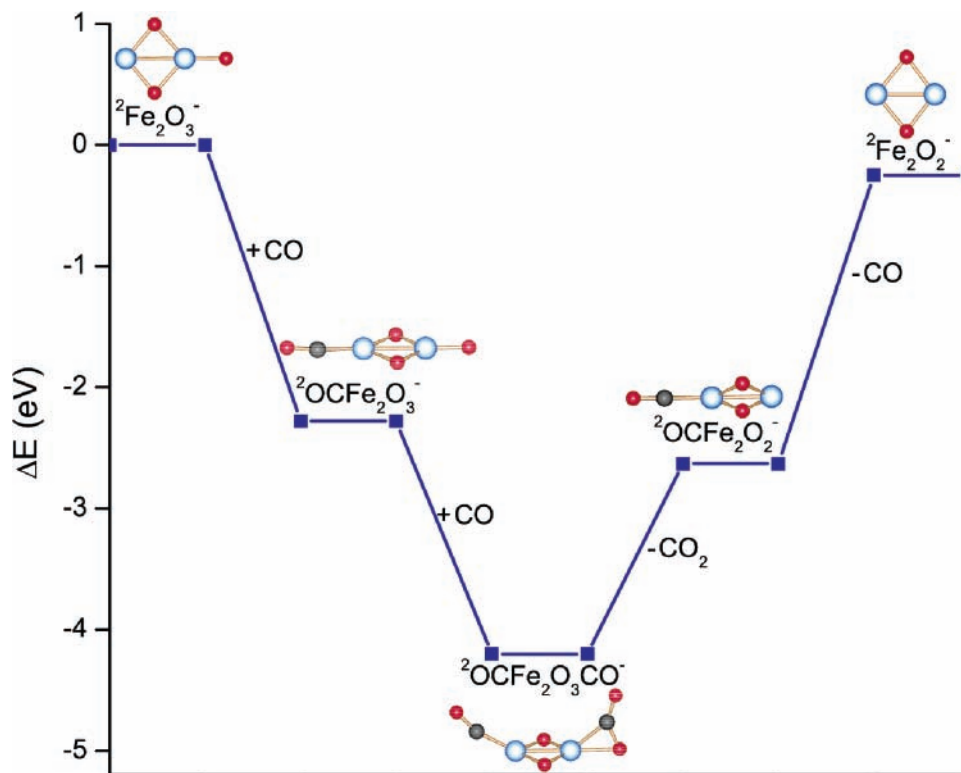


Figure 9. The change in energy (ΔE) for the reaction pathway of Fe_2O_3^- with CO. The superscripts indicate the spin multiplicity.

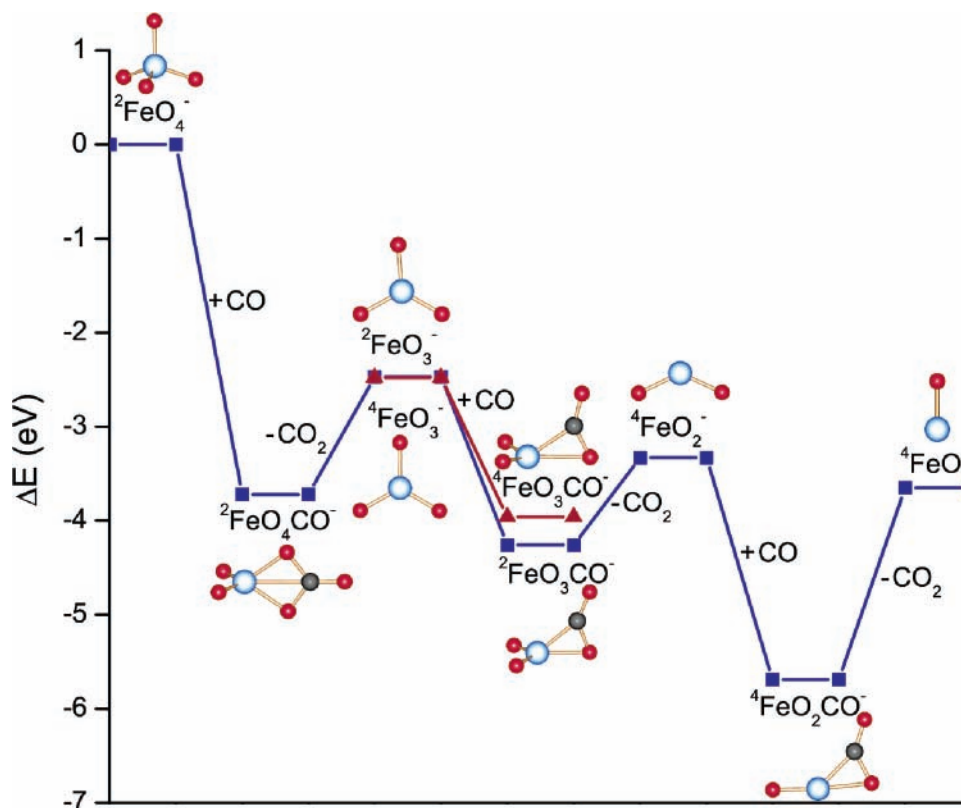


Figure 10. The change in energy (ΔE) for the reaction pathway of FeO_4^- with CO. The superscripts indicate the spin multiplicity.

the presence of $\text{Fe}_x\text{O}_{y-1}^-$ products suggests that oxygen is transferred from iron oxide clusters to CO to produce neutral CO_2 .

Figure 8 shows how the FeO_2^- and Fe_2O_3^- signals change as a function of increasing CO pressure. These cluster stoichiometries are the most efficient iron oxide anions to effect the CO oxidation reaction channel. Notice that both clusters are

composed of one more oxygen atom than the number of iron atoms. The energy calculated to remove an oxygen atom from these two clusters is approximately 6 eV. Therefore, the total energy needed is more than can be supplied through thermal collisions alone. This is supported by the fact that no oxygen atom loss products are observed in nitrogen studies conducted under the same experimental conditions (see Table 1). Confir-

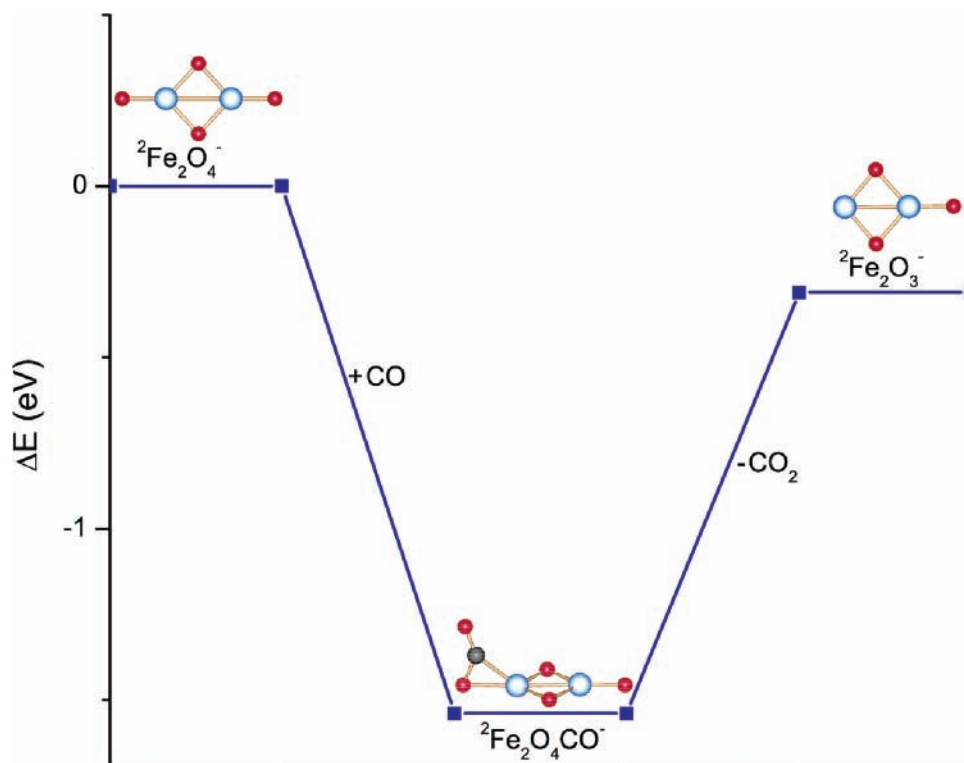


Figure 11. The change in energy (ΔE) for the reaction pathway of Fe_2O_4^- with CO. The superscripts indicate the spin multiplicity.

mation of the chemical reactivity of iron oxide anionic clusters with CO, therefore, can be seen by comparing the products to those observed with nitrogen.

We also calculated the reaction paths of the FeO_2^- , FeO_3^- , FeO_4^- , Fe_2O_3^- , and Fe_2O_4^- clusters with CO in order to understand the different reactivity they displayed. We found that for the most active clusters, FeO_2^- and Fe_2O_3^- , the reaction proceeds without barriers and follows a spin allowed path. The reaction path for Fe_2O_3^- is shown in Figure 9 with the exact values provided in supplemental Table S6. The first CO can attach to the Fe site or approach the O atom to form CO_2 . Our studies indicate that the more stable configuration corresponds to CO attached to the Fe site. This is consistent with the electron donating behavior of CO and the partial positive charge present on the Fe site, as revealed by the Mulliken population calculation. The subsequent CO then attaches to the O site forming CO_2 as shown in Figure 9. Previous studies on neutral Fe_2O_3 also obtained a similar reaction path.⁸

On the other hand, for FeO_4^- shown in Figure 10 and Table S7, the reaction with CO involves a change of spin which may slow the reaction. For the Fe_2O_4^- the reaction proceeds without barriers and follows a spin allowed path as revealed in Figure 11 and Table S8. In the initial step of the reaction, there is a gain in energy of only 1.54 eV for the absorption of the CO molecule compared with a gain of 4.20 eV for the absorption of two CO molecules in the Fe_2O_3^- cluster. This is because one of the Fe sites in Fe_2O_3^- is coordinated to only two O atoms and the first CO binds to this metal site. This difference in adsorption energy explains the difference in reactivity observed in the two clusters. It is important to note that the reaction pathways involve intermediate species and consequently require structural rearrangements. For small clusters, the reduced size is amenable to these rearrangements and consequently, the reaction barriers are far less than for the case of bulk surfaces where the extended geometry is less flexible toward structural changes.

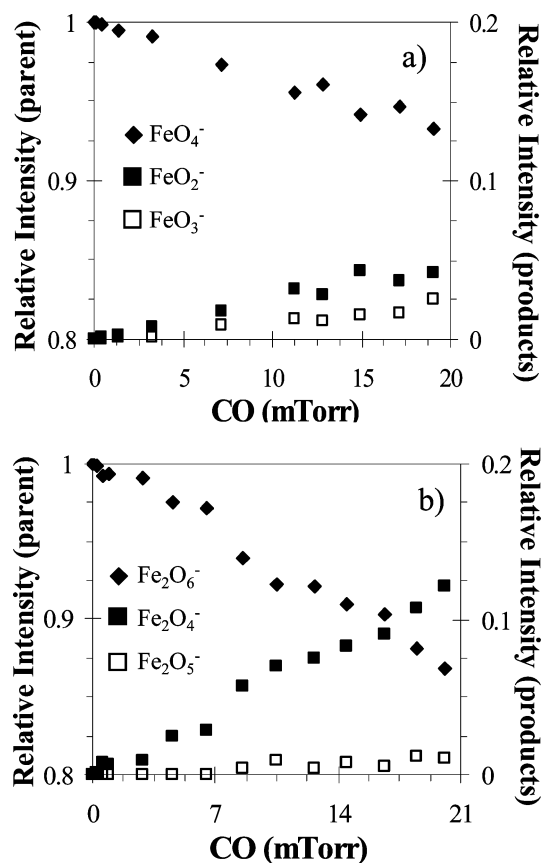


Figure 12. Relative intensities for (a) FeO_4^- and products species and (b) Fe_2O_6^- and product species as functions of increasing CO pressure. Notice both clusters show the dominant channel to be molecular oxygen loss over atomic oxygen loss.

Another reaction channel observed was the loss of molecular oxygen from FeO_4^- and Fe_2O_6^- . The change in cluster intensity with increasing CO pressure is shown in Figure 12, with O_2

loss as the dominant reaction channel followed by minor O atom loss. In the case of FeO_4^- , it is presumed at the high CO pressures in which FeO_2^- is detected as a product, the loss of molecular oxygen can be attributed to collision excitation that promotes the cluster to the structure containing an intact molecular oxygen unit. The same will apply for the formation of the higher-energy isomer of Fe_2O_6^- with a molecular oxygen unit. Both Fe_2O_6^- and FeO_4^- showed O_2 loss in the presence of nitrogen as seen in Table 1. These products were observed at higher nitrogen pressure under conditions where multiple collisions are expected.

Conclusion

It is shown that the nature of fragments produced in collision induced dissociation is governed not only by the overall energetics but also by the nature and multiplicity of the intermediate states marking the fragmentation. For FeO_n^- clusters, the studies show that while the loss of atomic O is the favored channel at lower oxygen coverage, the loss of O_2 becomes the dominant pathway when more oxygen surrounds a single Fe atom. In the case of Fe_2O_n^- clusters, the lowest-energy fragmentation pathways involve the loss of Fe or FeO units except for Fe_2O_6^- , where the lowest-energy pathway is the loss of an O_2 molecule. Small anionic iron oxide clusters are shown to enable the oxidation of CO at near thermal energies. The most active and selective iron oxides were composed of one more oxygen atom than iron atom. The increased reactivity is attributed to two reasons. (1) The energy required to remove an O atom in these clusters is less than the gain in energy to form CO_2 making the oxidation thermodynamically feasible. (2) The reduced size (compared to bulk species) allows structural rearrangements that eliminate/reduce the high reaction barriers and make the oxidation kinetically possible. We are in the process of extending these studies to cationic clusters and larger sizes to examine the effect of charge state and size on the catalytic conversion.

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Supporting Information Available: Tables with energies for dissociation pathways corresponding to figures described in the text; figure for the dissociation pathway for Fe_2O_6^- and corresponding energies; tables of energies for reaction pathways of figures in the text. This material is available free of charge via the Internet at <http://pubs.acs.org>.

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